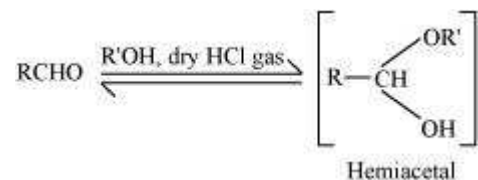


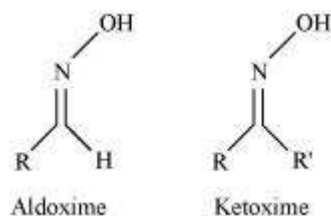
General structure of a hemiacetal

Aldehyde reacts with one molecule of a monohydric alcohol in the presence of dry HCl gas.

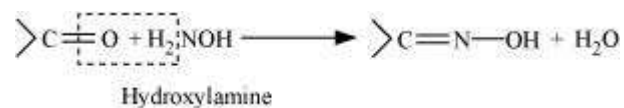


(vi) Oxime:

Oximes are a class of organic compounds having the general formula RR'CNOH , where R is an organic side chain and R' is either hydrogen or an organic side chain. If R' is H, then it is known as aldoxime and if R' is an organic side chain, it is known as ketoxime.

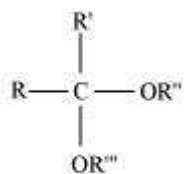


On treatment with hydroxylamine in a weakly acidic medium, aldehydes or ketones form oximes.



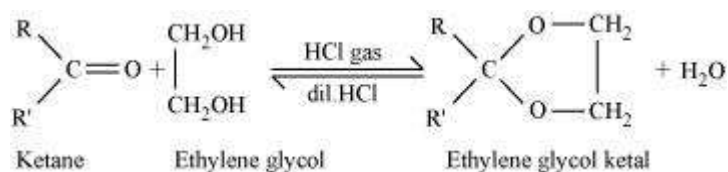
(vii) Ketal:

Ketals are gem-dialkoxyalkanes in which two alkoxy groups are present on the same carbon atom within the chain. The other two bonds of the carbon atom are connected to two alkyl groups.



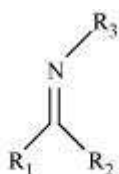
General structure of a ketal

Ketones react with ethylene glycol in the presence of dry HCl gas to give a cyclic product known as ethylene glycol ketals.



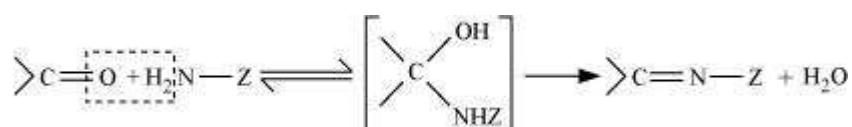
(viii) Imine:

Imines are chemical compounds containing a carbon nitrogen double bond.



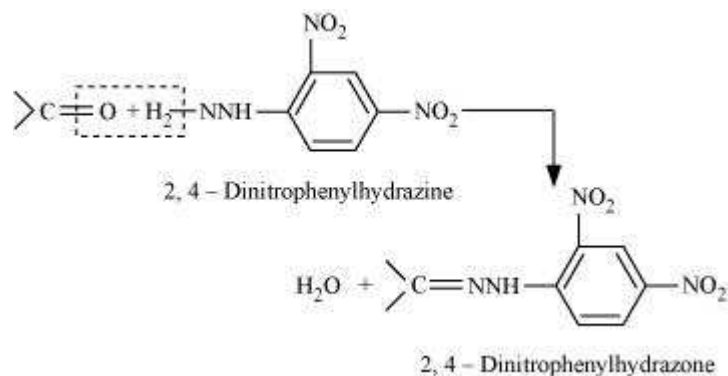
General structure of an imine

Imines are produced when aldehydes and ketones react with ammonia and its derivatives.



(ix) 2, 4-DNP-derivative:

2, 4-dinitrophenylhydrazones are 2, 4-DNP-derivatives, which are produced when aldehydes or ketones react with 2, 4-dinitrophenylhydrazine in a weakly acidic medium.

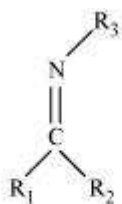


To identify and characterize aldehydes and ketones, 2, 4-DNP derivatives are used.

(x) Schiff's base:

Schiff's base (or azomethine) is a chemical compound containing a carbon-nitrogen double bond with the nitrogen atom connected to an aryl or alkyl group-but not hydrogen. They have the general formula $\text{R}_1\text{R}_2\text{C} = \text{NR}_3$. Hence, it is an imine.

It is named after a scientist, Hugo Schiff.



General structure of schiff's base

Aldehydes and ketones on treatment with primary aliphatic or aromatic amines in the presence of trace of an acid yields a Schiff's base.



Question 12.2:

Name the following compounds according to IUPAC system of nomenclature:

- (i) $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_2\text{CHO}$
- (ii) $\text{CH}_3\text{CH}_2\text{COCH}(\text{C}_2\text{H}_5)\text{CH}_2\text{CH}_2\text{Cl}$
- (iii) $\text{CH}_3\text{CH}=\text{CHCHO}$
- (iv) $\text{CH}_3\text{COCH}_2\text{COCH}_3$
- (v) $\text{CH}_3\text{CH}(\text{CH}_3)\text{CH}_2\text{C}(\text{CH}_3)_2\text{COCH}_3$
- (vi) $(\text{CH}_3)_3\text{CCH}_2\text{COOH}$
- (vii) $\text{OHCC}_6\text{H}_4\text{CHO}-p$

Answer

- (i) 4-methylpentanal
- (ii) 6-Chloro-4-ethylhexan-3-one
- (iii) But-2-en-1-al
- (iv) Pentane-2,4-dione
- (v) 3,3,5-Trimethylhexan-2-one
- (vi) 3,3-Dimethylbutanoic acid
- (vii) Benzene-1,4-dicarbaldehyde

Question 12.3:

Draw the structures of the following compounds.

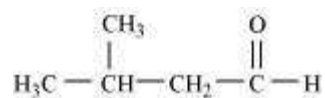
- (i) 3-Methylbutanal (ii) *p*-Nitropropiophenone
 - (iii) *p*-Methylbenzaldehyde (iv) 4-Methylpent-3-en-2-one
-

(v) 4-Chloropentan-2-one (vi) 3-Bromo-4-phenylpentanoic acid

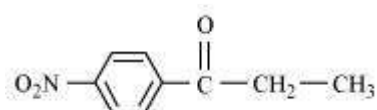
(vii) *p,p'*-Dihydroxybenzophenone (viii) Hex-2-en-4-ynoic acid

Answer

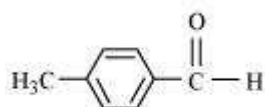
(i)



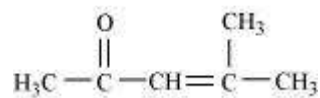
(ii)



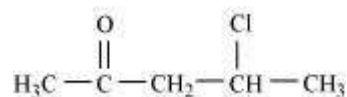
(iii)



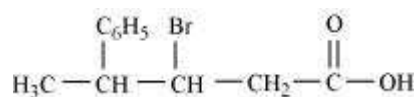
(iv)



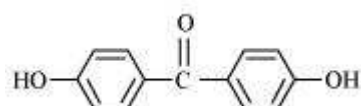
(v)



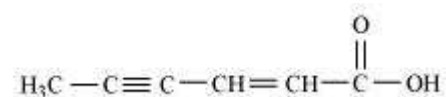
(vi)



(vii)

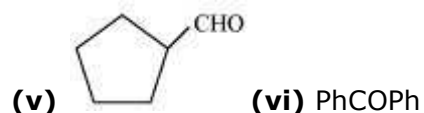
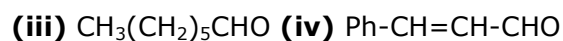
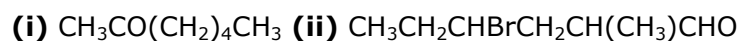


(viii)

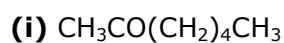


Question 12.4:

Write the IUPAC names of the following ketones and aldehydes. Wherever possible, give also common names.



Answer



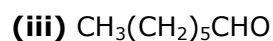
IUPAC name: Heptan-2-one

Common name: Methyl n-propyl ketone

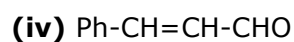


IUPAC name: 4-Bromo-2-methylhexanal

Common name: (γ -Bromo- α -methyl-caproaldehyde)



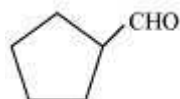
IUPAC name: Heptanal



IUPAC name: 3-phenylprop-2-enal

Common name: β -Phenylacrolein

(v)



IUPAC name: Cyclopentanecarbaldehyde

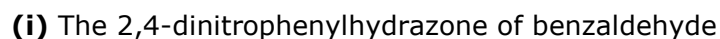


IUPAC name: Diphenylmethanone

Common name: Benzophenone

Question 12.5:

Draw structures of the following derivatives.

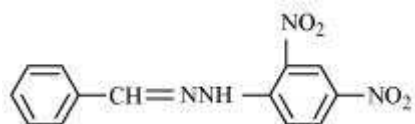


(v) The ethylene ketal of hexan-3-one

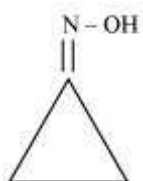
(vi) The methyl hemiacetal of formaldehyde

Answer

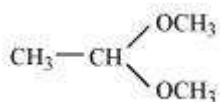
(i)



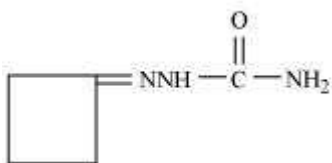
(ii)



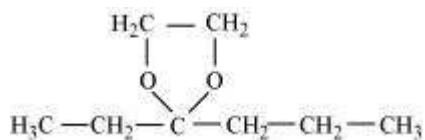
(iii)



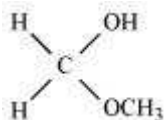
(iv)



(v)



(vi)



Question 12.6:

Predict the products formed when cyclohexanecarbaldehyde reacts with following reagents.

(i) PhMgBr and then H₃O⁺

(ii) Tollens' reagent

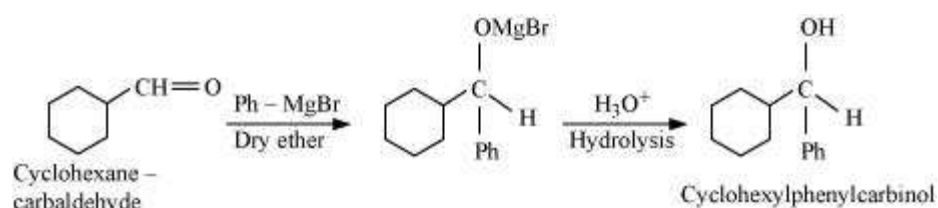
(iii) Semicarbazide and weak acid

(iv) Excess ethanol and acid

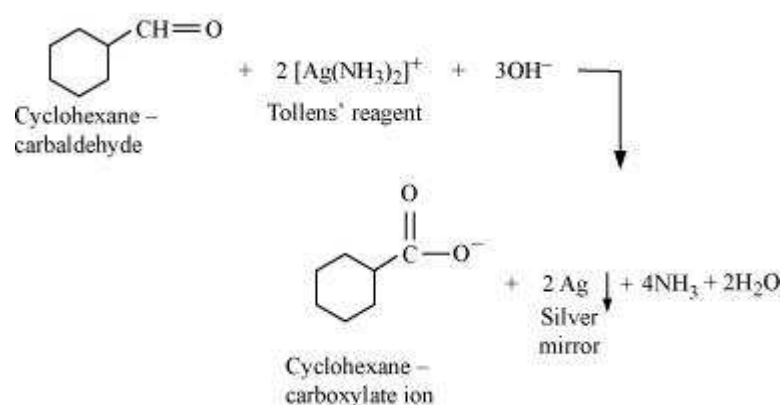
(v) Zinc amalgam and dilute hydrochloric acid

Answer

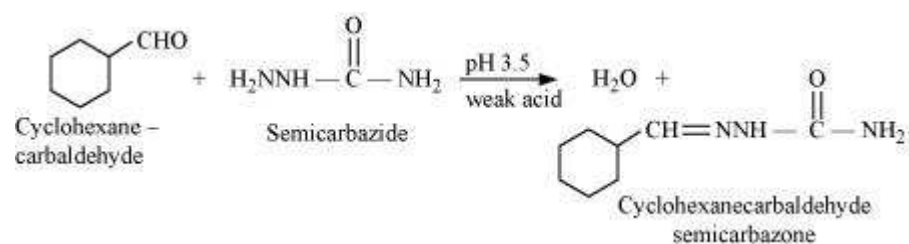
(i)



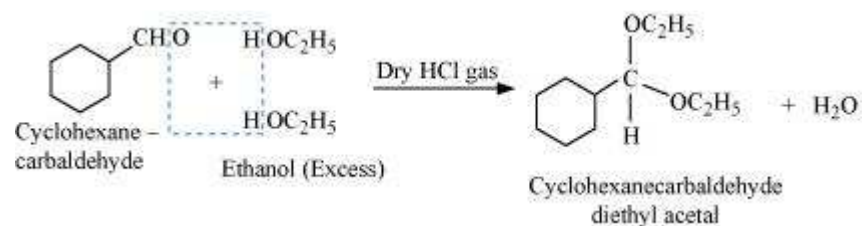
(ii)



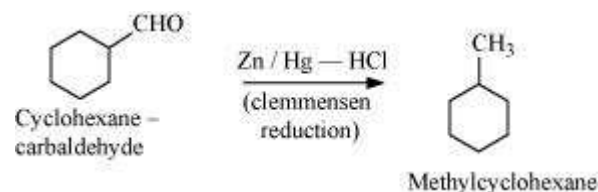
(iii)



(iv)



(v)



Question 12.7:

Which of the following compounds would undergo aldol condensation, which the Cannizzaro reaction and which neither? Write the structures of the expected products of aldol condensation and Cannizzaro reaction.

- (i) Methanal (ii) 2-Methylpentanal
(iii) Benzaldehyde (iv) Benzophenone
(v) Cyclohexanone (vi) 1-Phenylpropanone
(vii) Phenylacetaldehyde (viii) Butan-1-ol
(ix) 2, 2-Dimethylbutanal

Answer

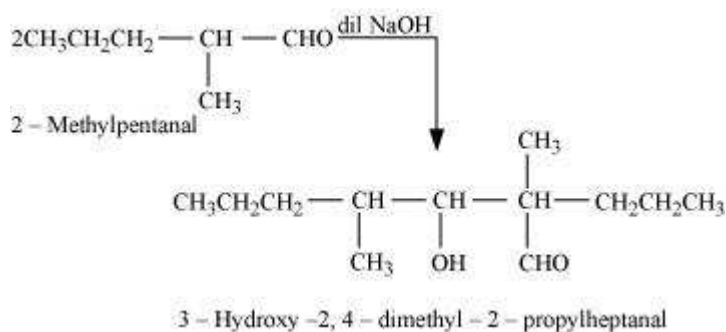
Aldehydes and ketones having at least one α -hydrogen undergo aldol condensation. The compounds (ii) 2-methylpentanal, (v) cyclohexanone, (vi) 1-phenylpropanone, and (vii) phenylacetaldehyde contain one or more α -hydrogen atoms. Therefore, these undergo aldol condensation.

Aldehydes having no α -hydrogen atoms undergo Cannizzaro reactions. The compounds (i) Methanal, (iii) Benzaldehyde, and (ix) 2, 2-dimethylbutanal do not have any α -hydrogen. Therefore, these undergo cannizzaro reactions.

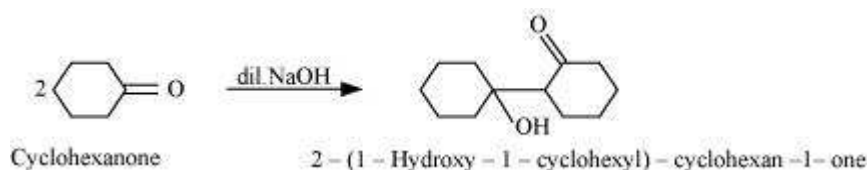
Compound (iv) Benzophenone is a ketone having no α -hydrogen atom and compound (viii) Butan-1-ol is an alcohol. Hence, these compounds do not undergo either aldol condensation or cannizzaro reactions.

Aldol condensation

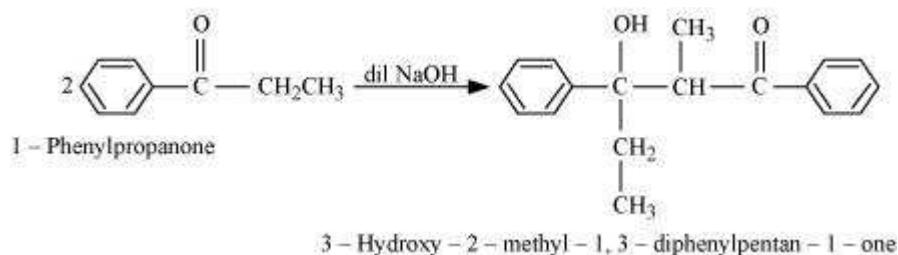
(ii)



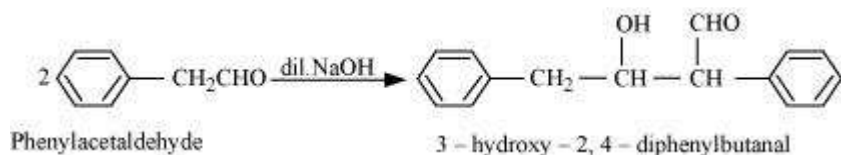
(v)



(vi)

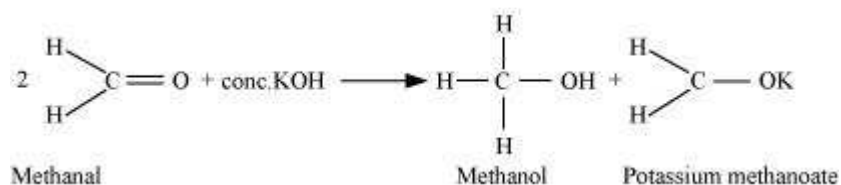


(vii)

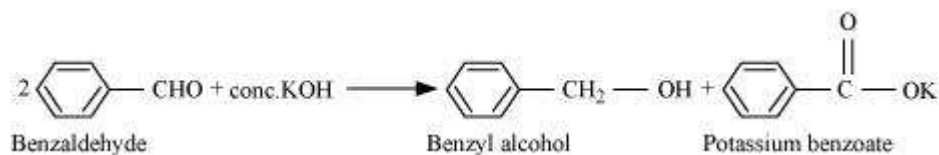


Cannizzaro reaction

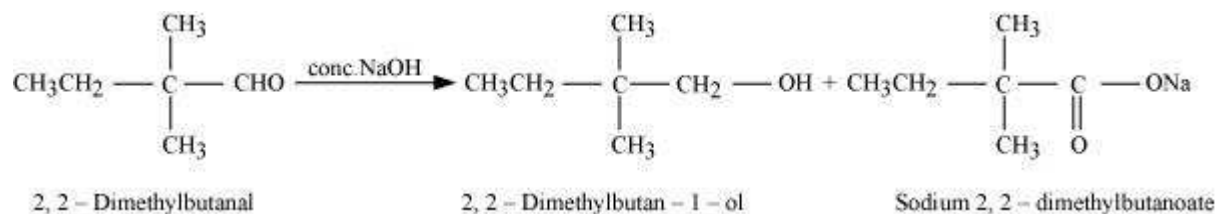
(i)



(iii)



(ix)



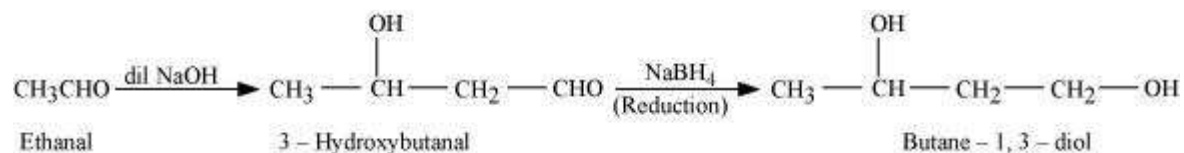
Question 12.8:

How will you convert ethanal into the following compounds?

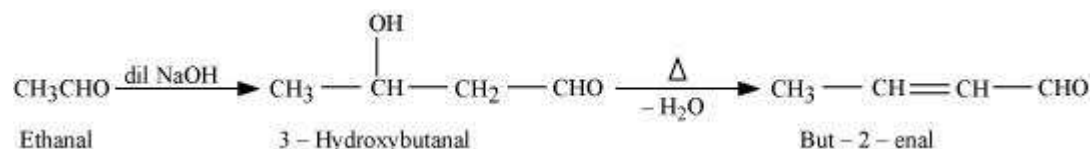
(i) Butane-1, 3-diol (ii) But-2-enal (iii) But-2-enoic acid

Answer

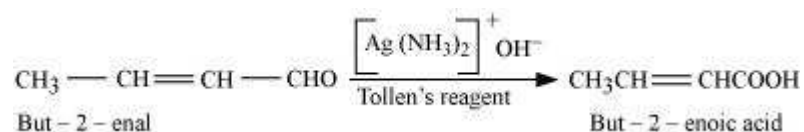
(i) On treatment with dilute alkali, ethanal produces 3-hydroxybutanal gives butane-1, 3-diol on reduction.



(ii) On treatment with dilute alkali, ethanal gives 3-hydroxybutanal which on heating produces but-2-enal.



(iii) When treated with Tollen's reagent, But-2-enal produced in the above reaction produces but-2-enoic acid .

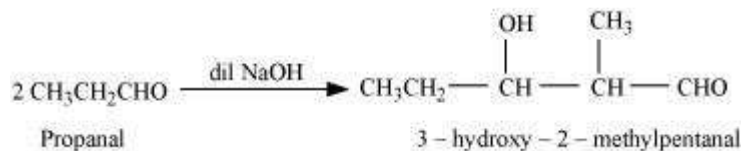


Question 12.9:

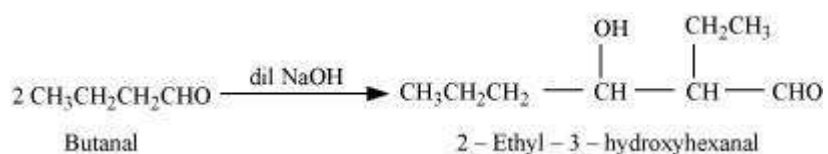
Write structural formulas and names of four possible aldol condensation products from propanal and butanal. In each case, indicate which aldehyde acts as nucleophile and which as electrophile.

Answer

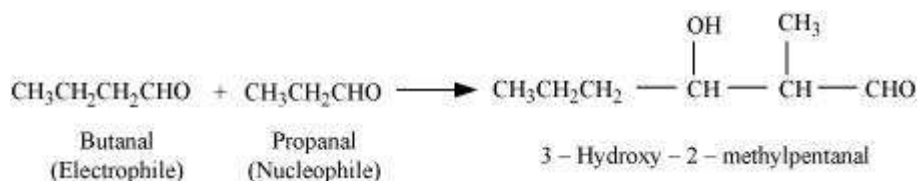
(i) Taking two molecules of propanal, one which acts as a nucleophile and the other as an electrophile.



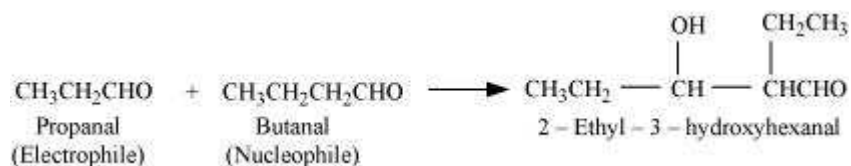
(ii) Taking two molecules of butanal, one which acts as a nucleophile and the other as an electrophile.



(iii) Taking one molecule each of propanal and butanal in which propanal acts as a nucleophile and butanal acts as an electrophile.



(iv) Taking one molecule each of propanal and butanal in which propanal acts as an electrophile and butanal acts as a nucleophile.



Question 12.10:

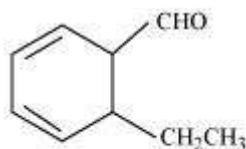
An organic compound with the molecular formula $\text{C}_9\text{H}_{10}\text{O}$ forms 2, 4-DNP derivative, reduces Tollens' reagent and undergoes Cannizzaro reaction. On vigorous oxidation, it gives 1, 2-benzenedicarboxylic acid. Identify the compound.

Answer

It is given that the compound (with molecular formula $\text{C}_9\text{H}_{10}\text{O}$) forms 2, 4-DNP derivative and reduces Tollen's reagent. Therefore, the given compound must be an aldehyde.

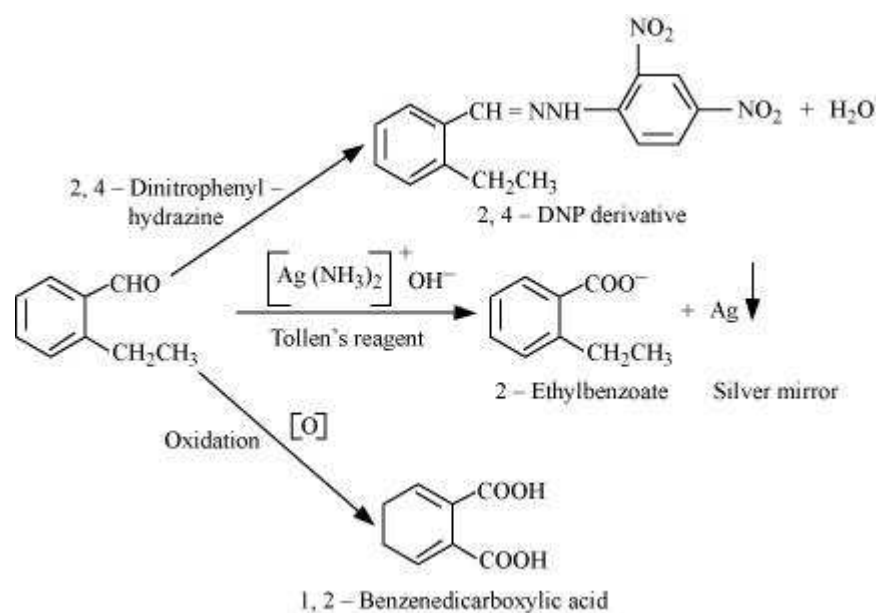
Again, the compound undergoes cannizzaro reaction and on oxidation gives 1, 2-benzenedicarboxylic acid. Therefore, the $-\text{CHO}$ group is directly attached to a benzene

ring and this benzaldehyde is ortho-substituted. Hence, the compound is 2-ethylbenzaldehyde.



2 – Ethylbenzaldehyde

The given reactions can be explained by the following equations.



Question 12.11:

An organic compound (A) (molecular formula $C_8H_{16}O_2$) was hydrolysed with dilute sulphuric acid to give a carboxylic acid (B) and an alcohol (C). Oxidation of (C) with chromic acid produced (B). (C) on dehydration gives but-1-ene. Write equations for the reactions involved.

Answer

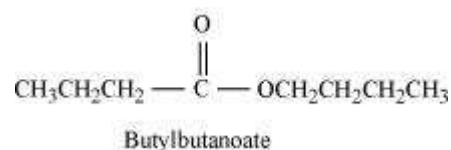
An organic compound A with molecular formula $C_8H_{16}O_2$ gives a carboxylic acid (B) and an alcohol (C) on hydrolysis with dilute sulphuric acid. Thus, compound A must be an ester. Further, alcohol C gives acid B on oxidation with chromic acid. Thus, B and C must contain equal number of carbon atoms.

Since compound A contains a total of 8 carbon atoms, each of B and C contain 4 carbon atoms.

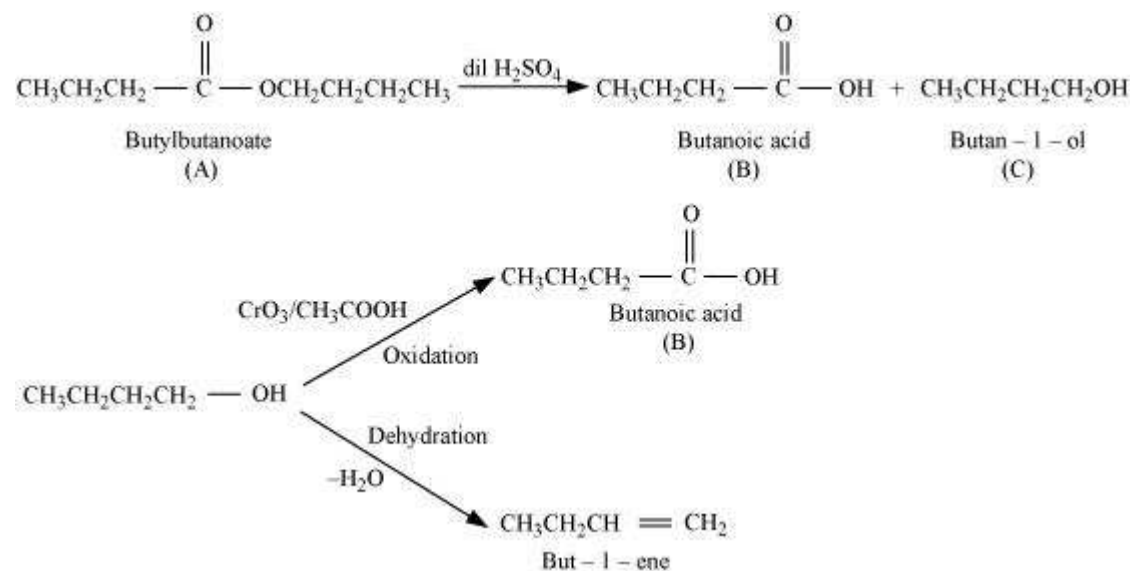
Again, on dehydration, alcohol C gives but-1-ene. Therefore, C is of straight chain and hence, it is butan-1-ol.

On oxidation, Butan-1-ol gives butanoic acid. Hence, acid B is butanoic acid.

Hence, the ester with molecular formula $C_8H_{16}O_2$ is butylbutanoate.



All the given reactions can be explained by the following equations.



Question 12.12:

Arrange the following compounds in increasing order of their property as indicated:

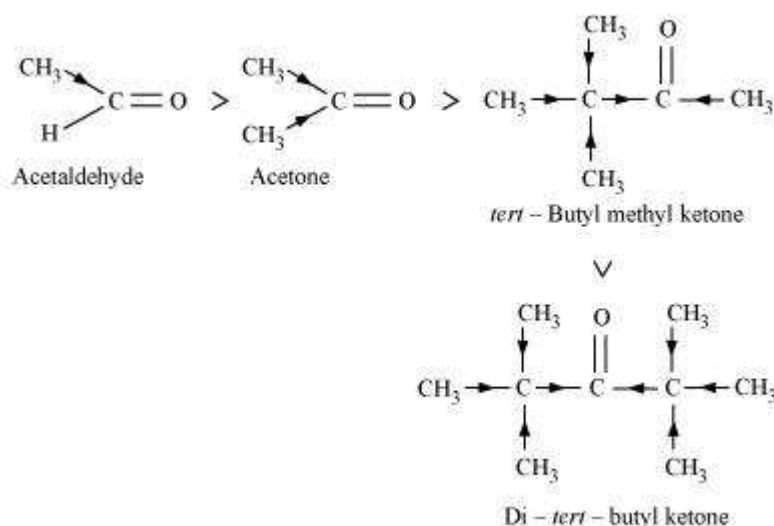
(i) Acetaldehyde, Acetone, Di-*tert*-butyl ketone, Methyl *tert*-butyl ketone (reactivity towards HCN)

(ii) $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$, $\text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH}$, $(\text{CH}_3)_2\text{CHCOOH}$, $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ (acid strength)

(iii) Benzoic acid, 4-Nitrobenzoic acid, 3,4-Dinitrobenzoic acid, 4-Methoxybenzoic acid (acid strength)

Answer

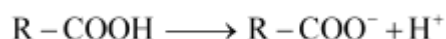
(i) When HCN reacts with a compound, the attacking species is a nucleophile, CN^- . Therefore, as the negative charge on the compound increases, its reactivity with HCN decreases. In the given compounds, the +I effect increases as shown below. It can be observed that steric hindrance also increases in the same



Hence, the given compounds can be arranged according to their increasing reactivities toward HCN as:

Di-*tert*-butyl ketone < Methyl *tert*-butyl ketone < Acetone < Acetaldehyde

(ii) After losing a proton, carboxylic acids gain a negative charge as shown:



Now, any group that will help stabilise the negative charge will increase the stability of the carboxyl ion and as a result, will increase the strength of the acid. Thus, groups having +I effect will decrease the strength of the acids and groups having -I effect will increase the strength of the acids. In the given compounds, $-\text{CH}_3$ group has +I effect and Br^- group has -I effect. Thus, acids containing Br^- are stronger.

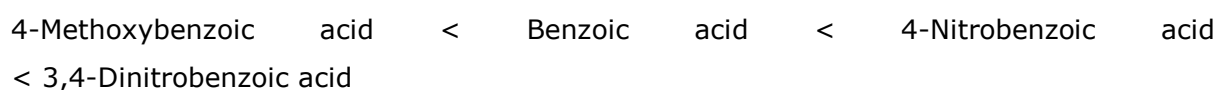
Now, the +I effect of isopropyl group is more than that of *n*-propyl group. Hence, $(\text{CH}_3)_2\text{CHCOOH}$ is a weaker acid than $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$.

Also, the -I effect grows weaker as distance increases. Hence, $\text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH}$ is a weaker acid than $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$.

Hence, the strengths of the given acids increase as:

$(\text{CH}_3)_2\text{CHCOOH} < \text{CH}_3\text{CH}_2\text{CH}_2\text{COOH} < \text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH} < \text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$

(iii) As we have seen in the previous case, electron-donating groups decrease the strengths of acids, while electron-withdrawing groups increase the strengths of acids. As methoxy group is an electron-donating group, 4-methoxybenzoic acid is a weaker acid than benzoic acid. Nitro group is an electron-withdrawing group and will increase the strengths of acids. As 3,4-dinitrobenzoic acid contains two nitro groups, it is a slightly stronger acid than 4-nitrobenzoic acid. Hence, the strengths of the given acids increase as:



Question 12.13:

Give simple chemical tests to distinguish between the following pairs of compounds.

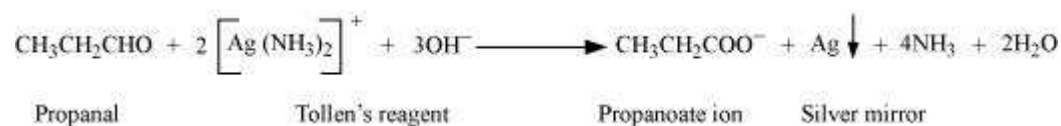
- (i)** Propanal and Propanone
- (ii)** Acetophenone and Benzophenone
- (iii)** Phenol and Benzoic acid
- (iv)** Benzoic acid and Ethyl benzoate
- (v)** Pentan-2-one and Pentan-3-one
- (vi)** Benzaldehyde and Acetophenone
- (vii)** Ethanal and Propanal

Answer

(i) Propanal and propanone can be distinguished by the following tests.

(a) Tollen's test

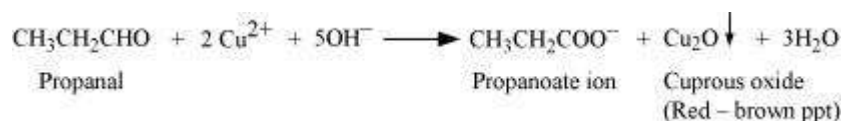
Propanal is an aldehyde. Thus, it reduces Tollen's reagent. But, propanone being a ketone does not reduce Tollen's reagent.



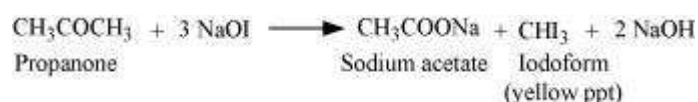
(b) Fehling's test

Aldehydes respond to Fehling's test, but ketones do not.

Propanal being an aldehyde reduces Fehling's solution to a red-brown precipitate of Cu_2O , but propanone being a ketone does not.

**(c) Iodoform test:**

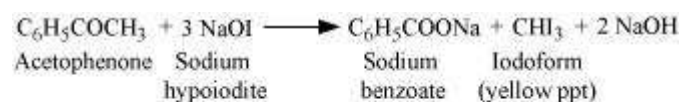
Aldehydes and ketones having at least one methyl group linked to the carbonyl carbon atom respond to iodoform test. They are oxidized by sodium hypiodite (NaOI) to give iodoforms. Propanone being a methyl ketone responds to this test, but propanal does not.



(ii) Acetophenone and Benzophenone can be distinguished using the iodoform test.

Iodoform test:

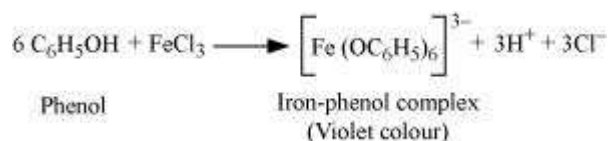
Methyl ketones are oxidized by sodium hypiodite to give yellow ppt. of iodoform. Acetophenone being a methyl ketone responds to this test, but benzophenone does not.



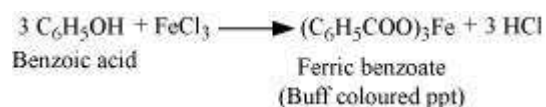
(iii) Phenol and benzoic acid can be distinguished by ferric chloride test.

Ferric chloride test:

Phenol reacts with neutral FeCl_3 to form an iron-phenol complex giving violet colouration.



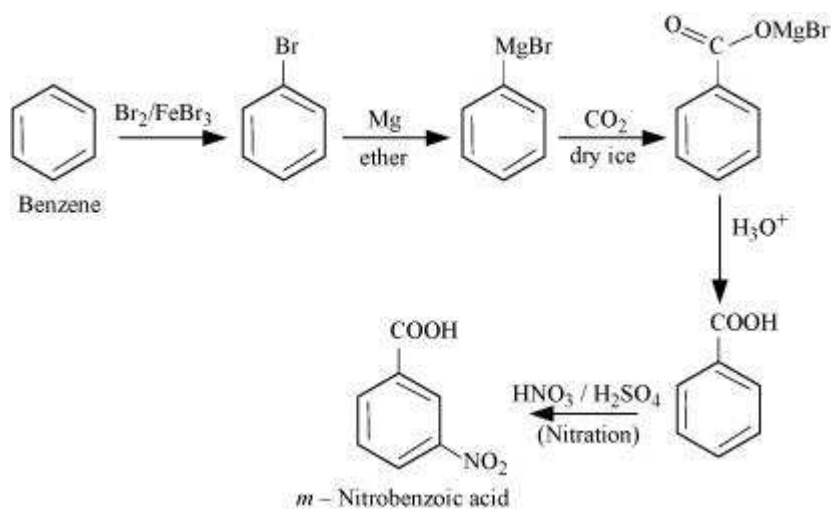
But benzoic acid reacts with neutral FeCl_3 to give a buff coloured ppt. of ferric benzoate.



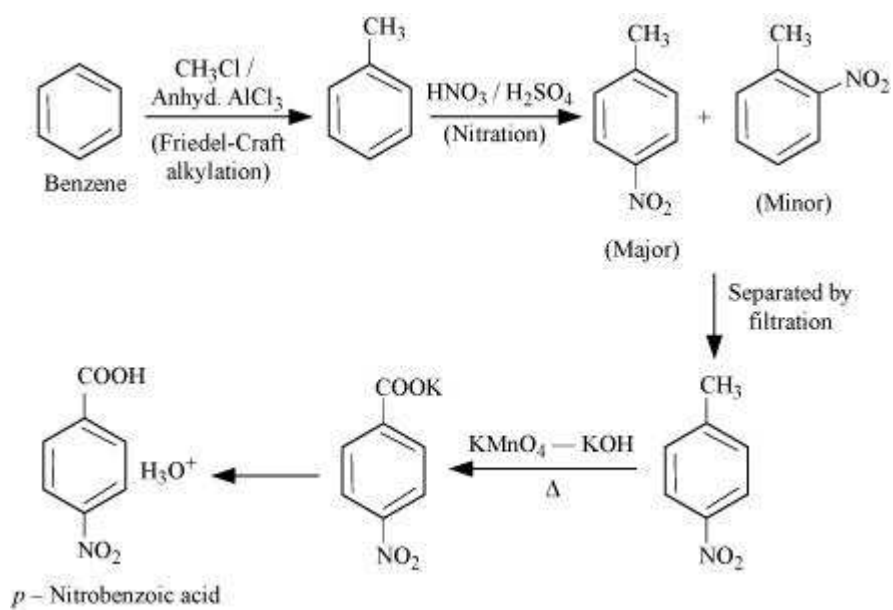
(iv) Benzoic acid and Ethyl benzoate can be distinguished by sodium bicarbonate test.

Sodium bicarbonate test:

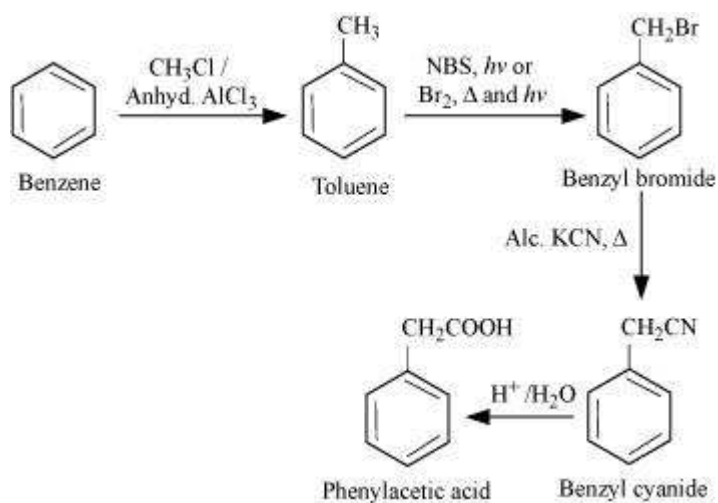
Acids react with NaHCO_3 to produce brisk effervescence due to the evolution of CO_2 gas. Benzoic acid being an acid responds to this test, but ethylbenzoate does not.



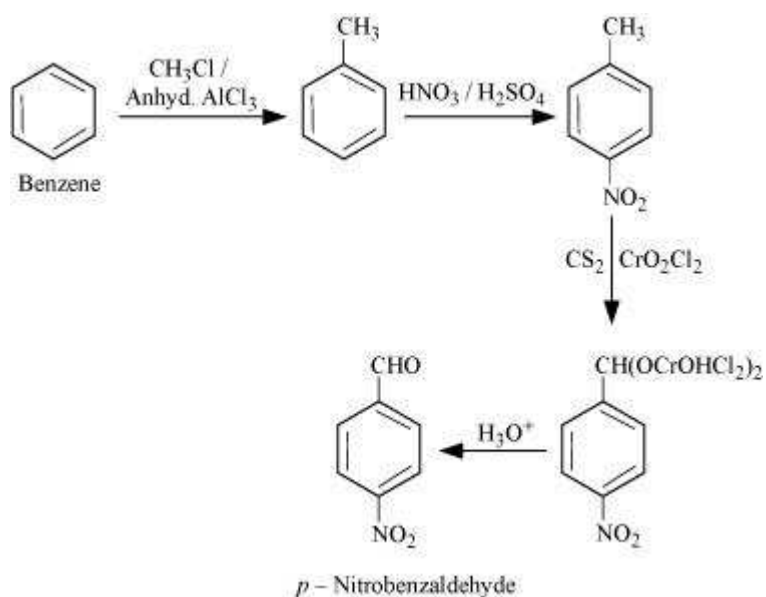
(iii)



(iv)



(v)



Question 12.15:

How will you bring about the following conversions in not more than two steps?

- (i) Propanone to Propene
- (ii) Benzoic acid to Benzaldehyde
- (iii) Ethanol to 3-Hydroxybutanal
- (iv) Benzene to *m*-Nitroacetophenone
- (v) Benzaldehyde to Benzophenone
- (vi) Bromobenzene to 1-Phenylethanol

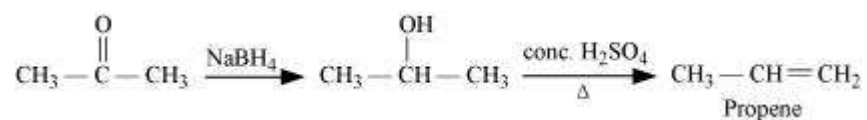
(vii) Benzaldehyde to 3-Phenylpropan-1-ol

(viii) Benzaldehyde to α -Hydroxyphenylacetic acid

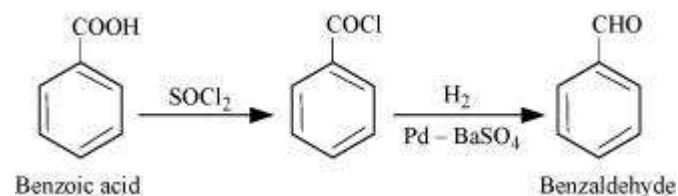
(ix) Benzoic acid to *m*- Nitrobenzyl alcohol

Answer

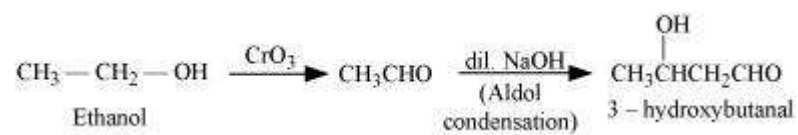
(i)



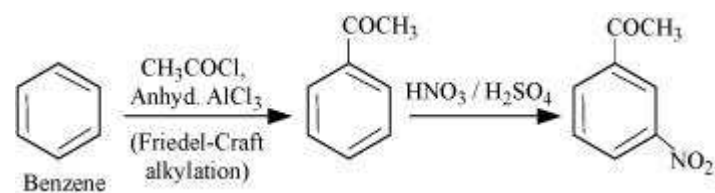
(ii)



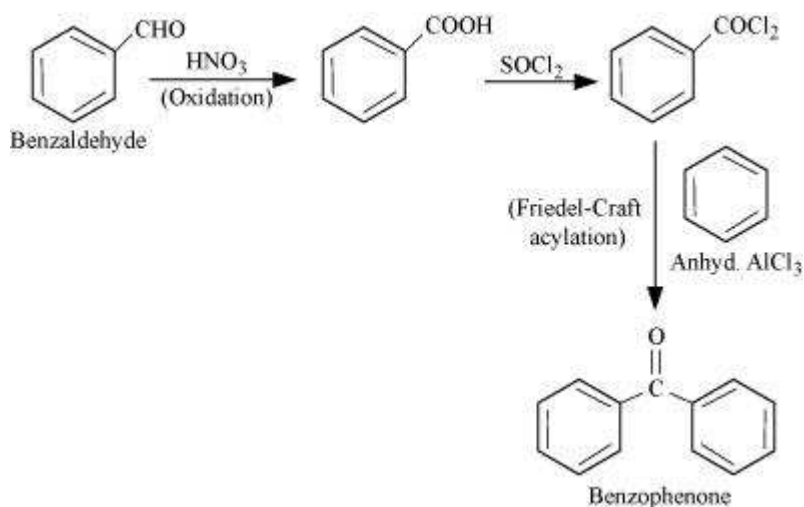
(iii)



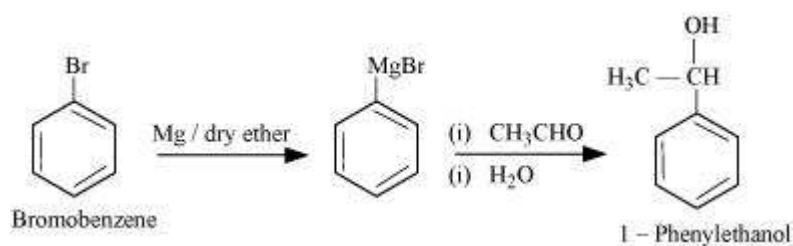
(iv)



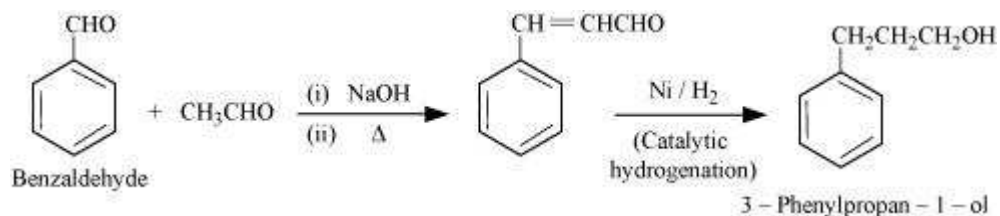
(v)



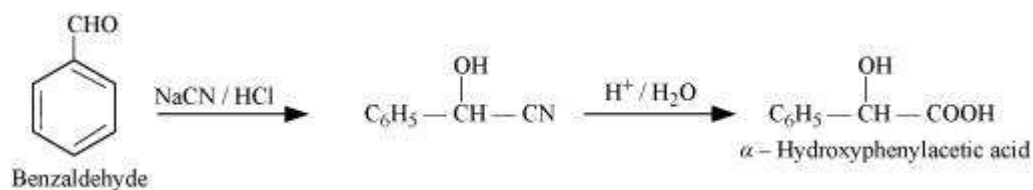
(vi)



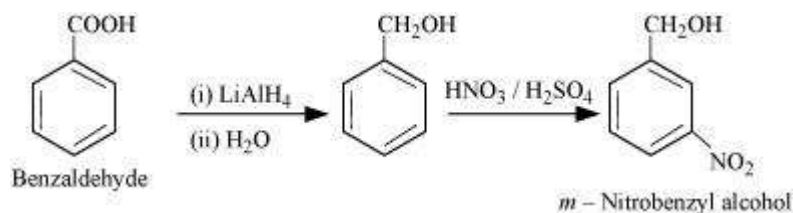
(vii)



(viii)



(ix)



Question 12.16:

Describe the following:

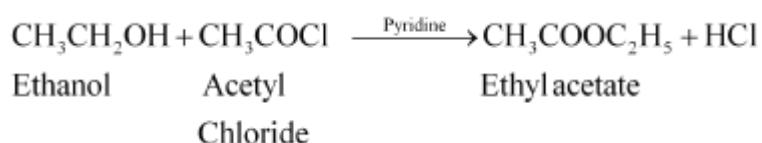
- (i) Acetylation
- (ii) Cannizzaro reaction
- (iii) Cross aldol condensation
- (iv) Decarboxylation

Answer

(i) Acetylation

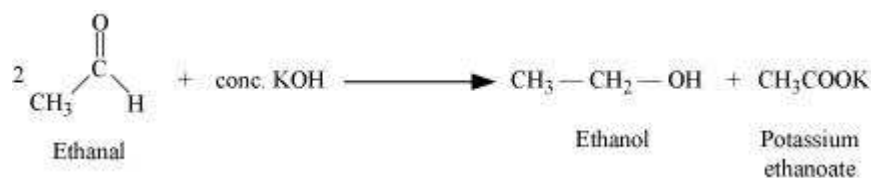
The introduction of an acetyl functional group into an organic compound is known as acetylation. It is usually carried out in the presence of a base such as pyridine, dimethylaniline, etc. This process involves the substitution of an acetyl group for an active hydrogen atom. Acetyl chloride and acetic anhydride are commonly used as acetylating agents.

For example, acetylation of ethanol produces ethyl acetate.

**(ii) Cannizzaro reaction:**

The self oxidation-reduction (disproportionation) reaction of aldehydes having no α -hydrogens on treatment with concentrated alkalis is known as the Cannizzaro reaction. In this reaction, two molecules of aldehydes participate where one is reduced to alcohol and the other is oxidized to carboxylic acid.

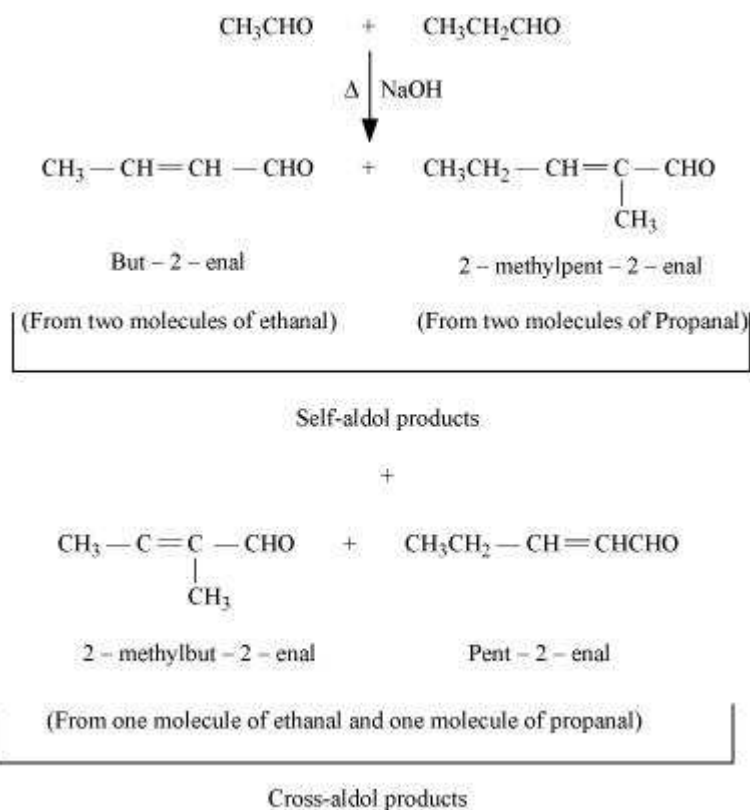
For example, when ethanol is treated with concentrated potassium hydroxide, ethanol and potassium ethanoate are produced.

**(iii) Cross-aldol condensation:**

When aldol condensation is carried out between two different aldehydes, or two different ketones, or an aldehyde and a ketone, then the reaction is called a cross-aldol

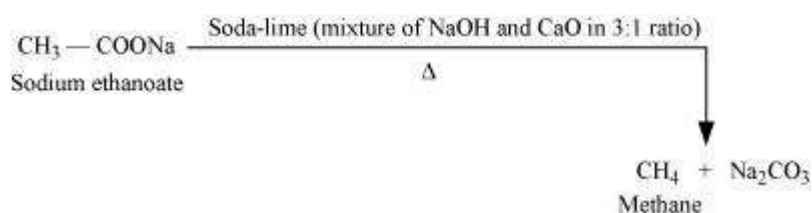
condensation. If both the reactants contain α -hydrogens, four compounds are obtained as products.

For example, ethanal and propanal react to give four products.



(iv) Decarboxylation:

Decarboxylation refers to the reaction in which carboxylic acids lose carbon dioxide to form hydrocarbons when their sodium salts are heated with soda-lime.

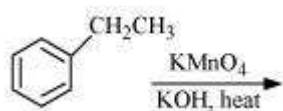


Decarboxylation also takes place when aqueous solutions of alkali metal salts of carboxylic acids are electrolyzed. This electrolytic process is known as Kolbe's electrolysis.

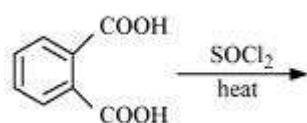
Question 12.17:

Complete each synthesis by giving missing starting material, reagent or products

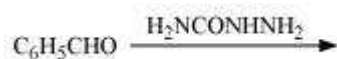
(i)



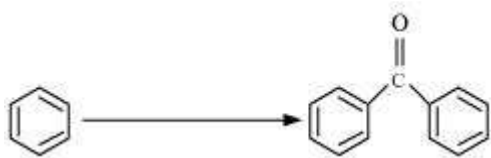
(ii)



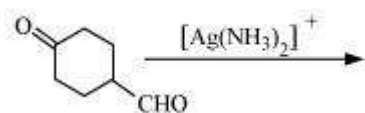
(iii)



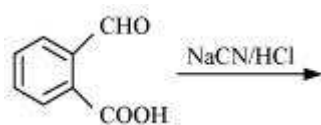
(iv)



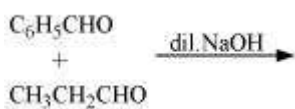
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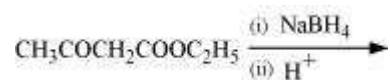
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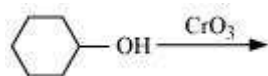
(vii)



(viii)



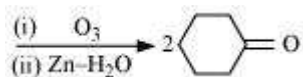
(ix)



(x)

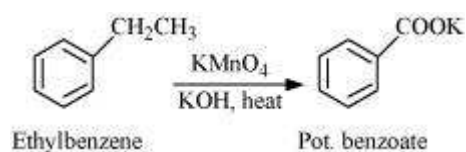


(xi)

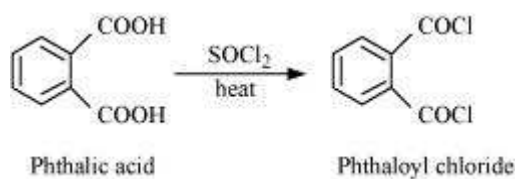


Answer

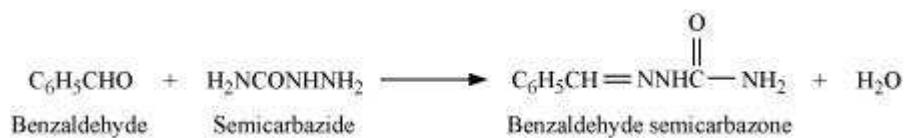
(i)



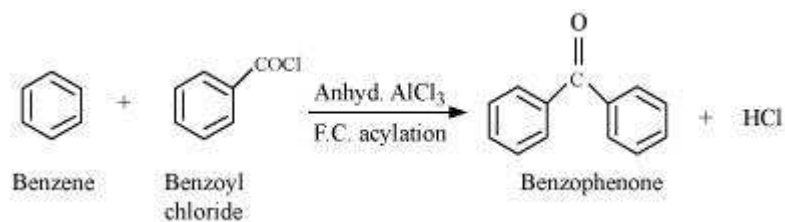
(ii)



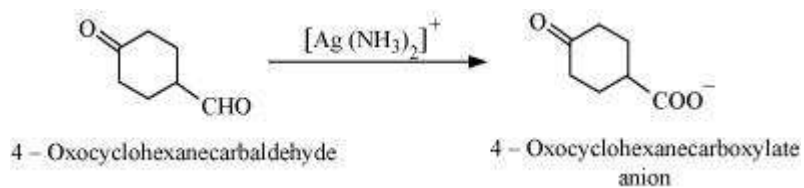
(iii)



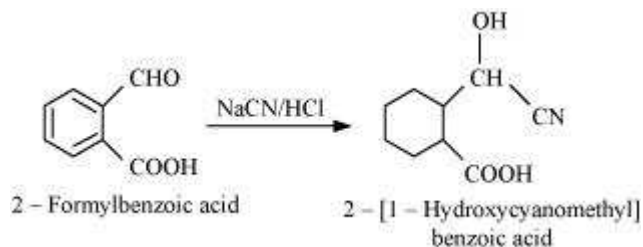
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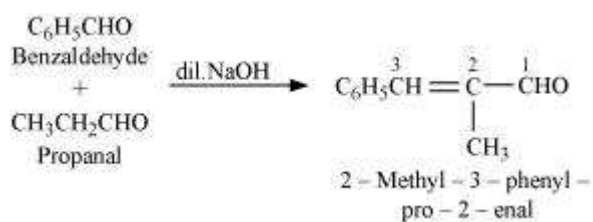
(v)



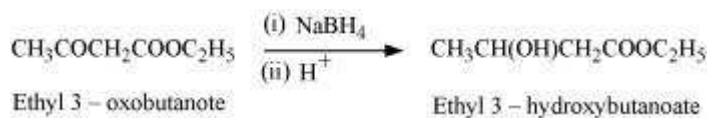
(vi)



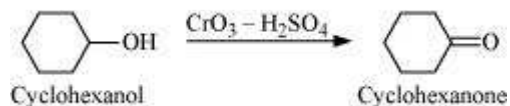
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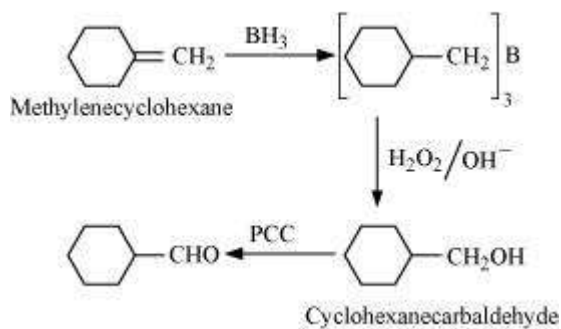
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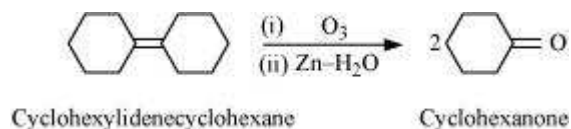
(ix)



(x)

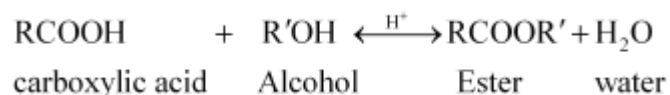


(xi)



Question 12.18:

(iii) Ester along with water is formed reversibly from a carboxylic acid and an alcohol in presence of an acid.



If either water or ester is not removed as soon as it is formed, then it reacts to give back the reactants as the reaction is reversible. Therefore, to shift the equilibrium in the forward direction i.e., to produce more ester, either of the two should be removed.

Question 12.19:

An organic compound contains 69.77% carbon, 11.63% hydrogen and rest oxygen. The molecular mass of the compound is 86. It does not reduce Tollens' reagent but forms an addition compound with sodium hydrogensulphite and give positive iodoform test. On vigorous oxidation it gives ethanoic and propanoic acid. Write the possible structure of the compound.

Answer

$$\% \text{ of carbon} = 69.77 \%$$

$$\% \text{ of hydrogen} = 11.63 \%$$

$$\begin{aligned} \% \text{ of oxygen} &= \{100 - (69.77 + 11.63)\} \% \\ &= 18.6 \% \end{aligned}$$

Thus, the ratio of the number of carbon, hydrogen, and oxygen atoms in the organic compound can be given as:

$$\text{C:H:O} = \frac{69.77}{12} : \frac{11.63}{1} : \frac{18.6}{16}$$

$$= 5.81:11.63:1.16$$

$$= 5:10:1$$

Therefore, the empirical formula of the compound is $\text{C}_5\text{H}_{10}\text{O}$. Now, the empirical formula mass of the compound can be given as:

$$\begin{aligned} 5 \times 12 + 10 \times 1 + 1 \times 16 \\ = 86 \end{aligned}$$

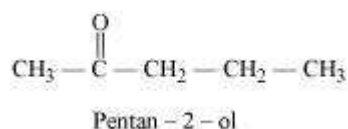
Molecular mass of the compound = 86

Therefore, the molecular formula of the compound is given by $\text{C}_5\text{H}_{10}\text{O}$.

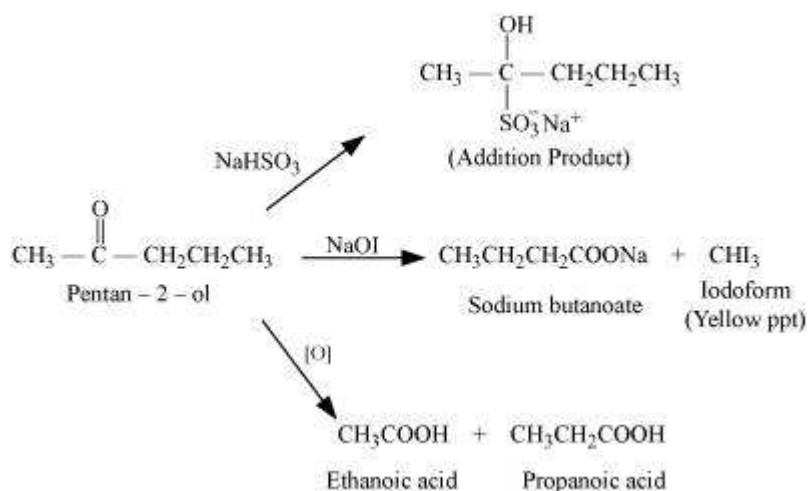
Since the given compound does not reduce Tollen's reagent, it is not an aldehyde. Again, the compound forms sodium hydrogen sulphate addition products and gives a positive iodoform test. Since the compound is not an aldehyde, it must be a methyl ketone.

The given compound also gives a mixture of ethanoic acid and propanoic acid.

Hence, the given compound is pentan-2-one.



The given reactions can be explained by the following equations:

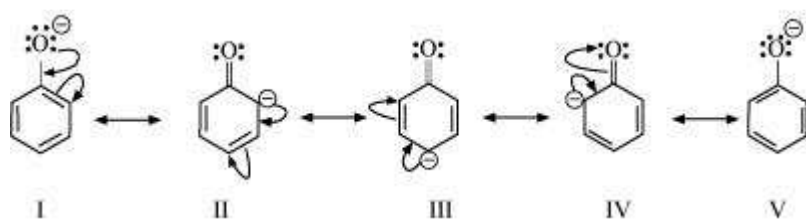


Question 12.20:

Although phenoxide ion has more number of resonating structures than carboxylate ion, carboxylic acid is a stronger acid than phenol. Why?

Answer

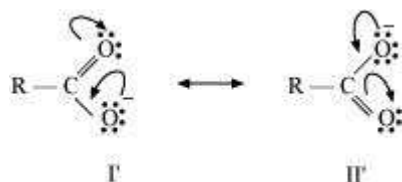
Resonance structures of phenoxide ion are:



It can be observed from the resonance structures of phenoxide ion that in II, III and IV, less electronegative carbon atoms carry a negative charge. Therefore, these three structures contribute negligibly towards the resonance stability of the phenoxide ion.

Hence, these structures can be eliminated. Only structures I and V carry a negative charge on the more electronegative oxygen atom.

Resonance structures of carboxylate ion are:



In the case of carboxylate ion, resonating structures I' and II' contain a charge carried by a more electronegative oxygen atom.

Further, in resonating structures I' and II', the negative charge is delocalized over two oxygen atoms. But in resonating structures I and V of the phenoxide ion, the negative charge is localized on the same oxygen atom. Therefore, the resonating structures of carboxylate ion contribute more towards its stability than those of phenoxide ion. As a result, carboxylate ion is more resonance-stabilized than phenoxide ion. Hence, carboxylic acid is a stronger acid than phenol.

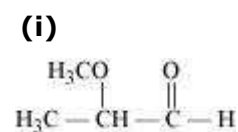
Text solution

Question 12.1:

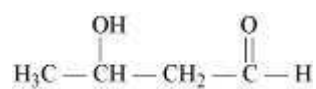
Write the structures of the following compounds.

- (i) α -Methoxypropionaldehyde
- (ii) 3-Hydroxybutanal
- (iii) 2-Hydroxycyclopentane carbaldehyde
- (iv) 4-Oxopentanal
- (v) Di-sec-butyl ketone
- (vi) 4-Fluoroacetophenone

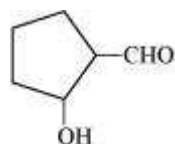
Answer



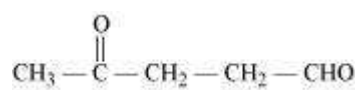
(ii)



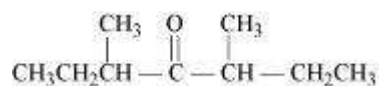
(iii)



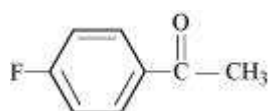
(iv)



(v)



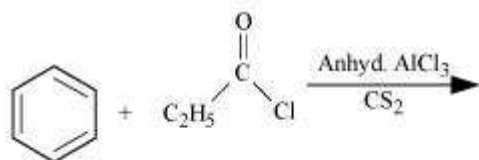
(vi)



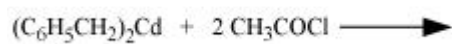
Question 12.2:

Write the structures of products of the following reactions;

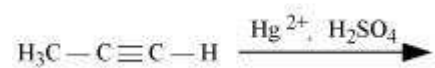
(i)



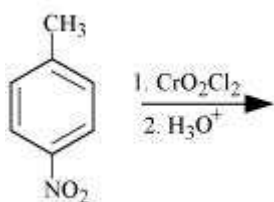
(ii)



(iii)

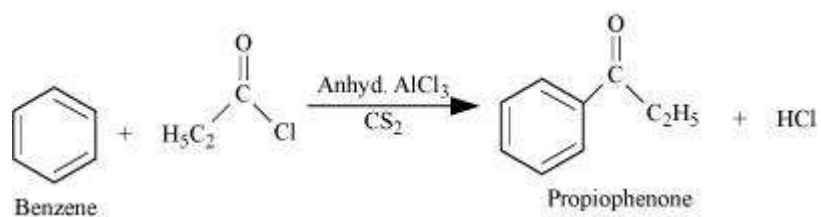


(iv)

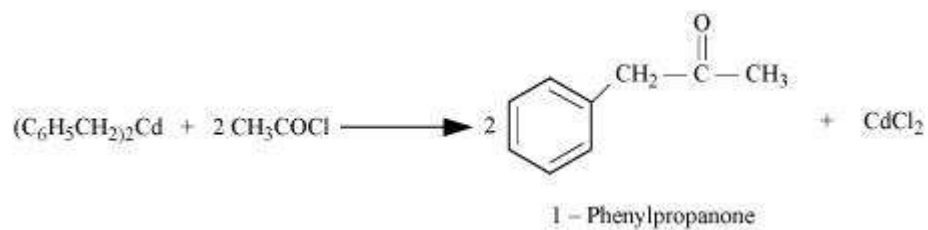


Answer

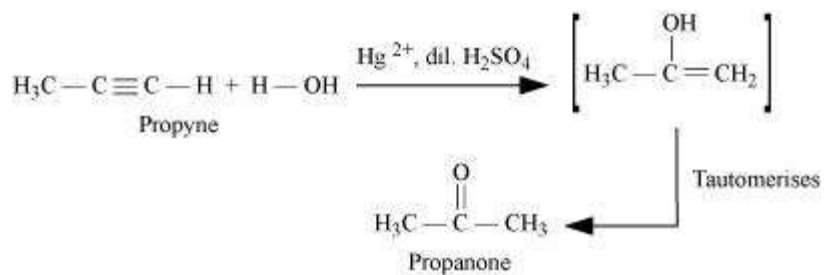
i.



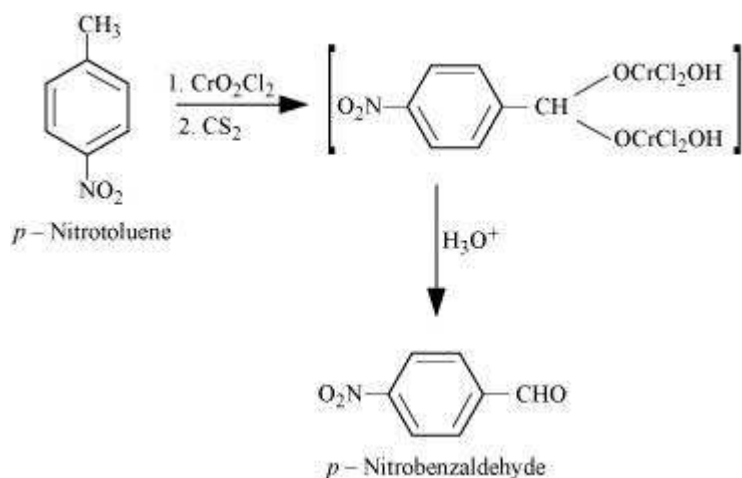
ii.



iii.



(iv)



Question 12.3:

Arrange the following compounds in increasing order of their boiling points.

CH_3CHO , $\text{CH}_3\text{CH}_2\text{OH}$, CH_3OCH_3 , $\text{CH}_3\text{CH}_2\text{CH}_3$

Answer

The molecular masses of the given compounds are in the range 44 to 46. $\text{CH}_3\text{CH}_2\text{OH}$ undergoes extensive intermolecular H-bonding, resulting in the association of molecules. Therefore, it has the highest boiling point. CH_3CHO is more polar than CH_3OCH_3 and so CH_3CHO has stronger intermolecular dipole – dipole attraction than CH_3OCH_3 . $\text{CH}_3\text{CH}_2\text{CH}_3$ has only weak van der Waals force. Thus, the arrangement of the given compounds in the increasing order of their boiling points is given by:

$\text{CH}_3\text{CH}_2\text{CH}_3 < \text{CH}_3\text{OCH}_3 < \text{CH}_3\text{CHO} < \text{CH}_3\text{CH}_2\text{OH}$

Question 12.4:

Arrange the following compounds in increasing order of their reactivity in nucleophilic addition reactions.

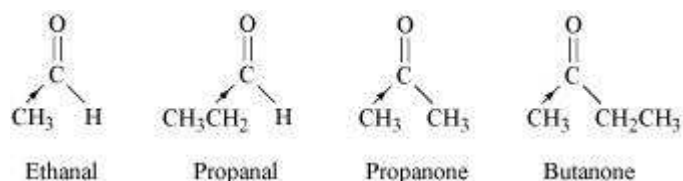
(i) Ethanal, Propanal, Propanone, Butanone.

(ii) Benzaldehyde, *p*-Tolualdehyde, *p*-Nitrobenzaldehyde, Acetophenone.

Hint: Consider steric effect and electronic effect.

Answer

(i)



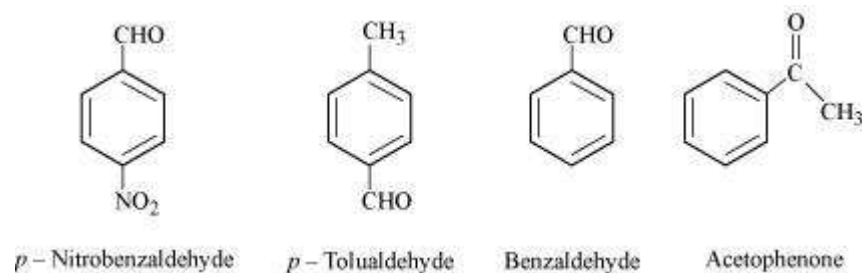
The +I effect of the alkyl group increases in the order:

Ethanal < Propanal < Propanone < Butanone

The electron density at the carbonyl carbon increases with the increase in the +I effect. As a result, the chances of attack by a nucleophile decrease. Hence, the increasing order of the reactivities of the given carbonyl compounds in nucleophilic addition reactions is:

Butanone < Propanone < Propanal < Ethanal

(ii)



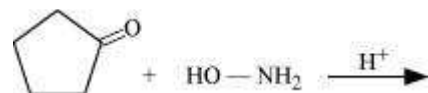
The +I effect is more in ketone than in aldehyde. Hence, acetophenone is the least reactive in nucleophilic addition reactions. Among aldehydes, the +I effect is the highest in *p*-tolualdehyde because of the presence of the electron-donating $-CH_3$ group and the lowest in *p*-nitrobenzaldehyde because of the presence of the electron-withdrawing $-NO_2$ group. Hence, the increasing order of the reactivities of the given compounds is:

Acetophenone < *p*-tolualdehyde < Benzaldehyde < *p*-Nitrobenzaldehyde

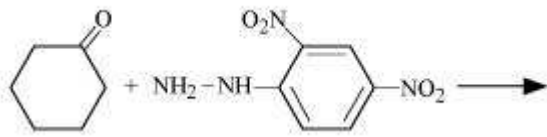
Question 12.5:

Predict the products of the following reactions:

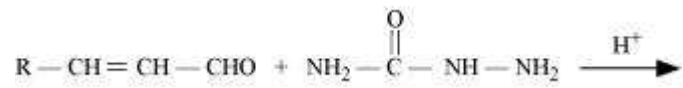
(i)



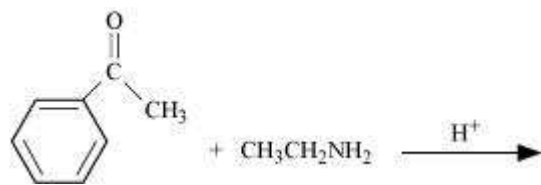
(ii)



(iii)

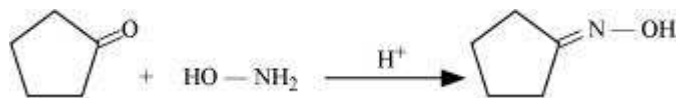


(iv)

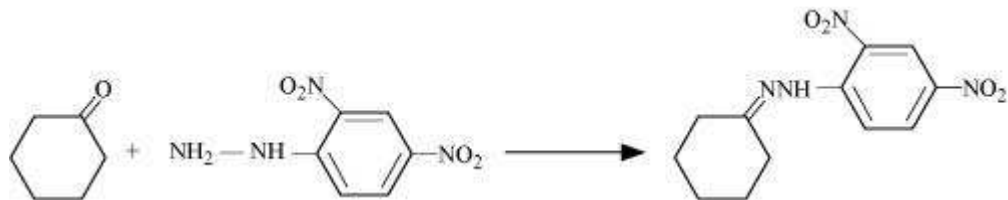


Answer

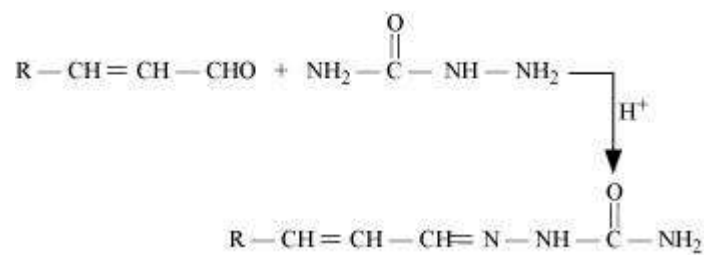
(i)



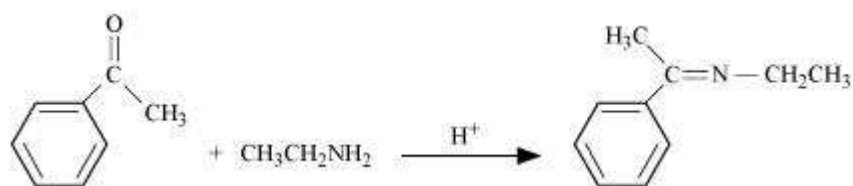
(ii)



(iii)



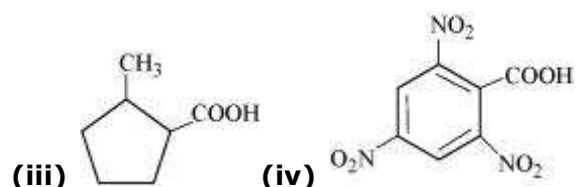
(iv)



Question 12.6:

Give the IUPAC names of the following compounds:

(i) $\text{PhCH}_2\text{CH}_2\text{COOH}$ **(ii)** $(\text{CH}_3)_2\text{C}=\text{CHCOOH}$



Answer

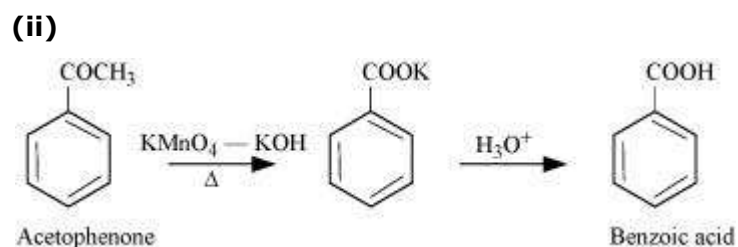
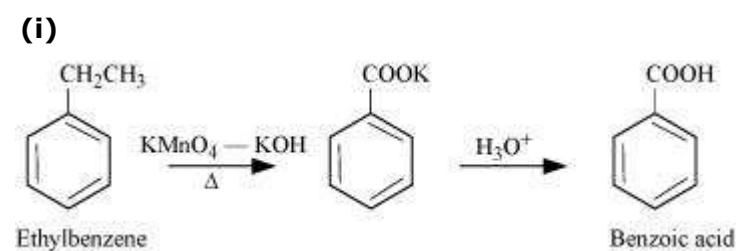
- (i)** 3-Phenylpropanoic acid
(ii) 3-Methylbut-2-enoic acid
(iii) 2-Methylcyclopentanecarboxylic acid
(iv) 2,4,6-Trinitrobenzoic acid

Question 12.7:

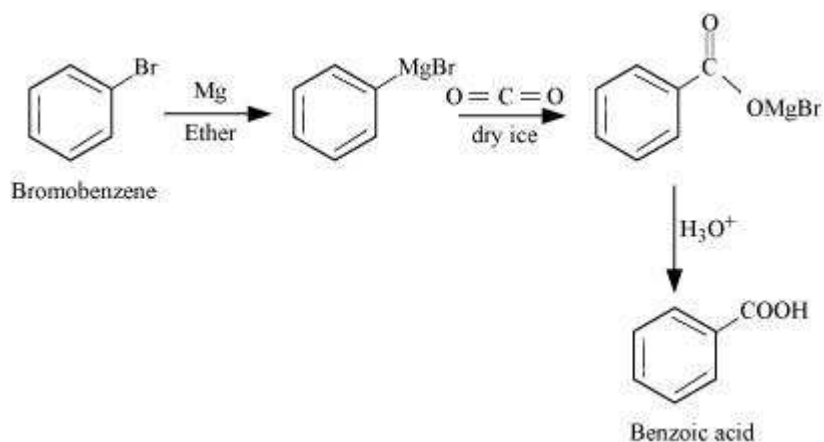
Show how each of the following compounds can be converted to benzoic acid.

- (i)** Ethylbenzene **(ii)** Acetophenone
(iii) Bromobenzene **(iv)** Phenylethene (Styrene)

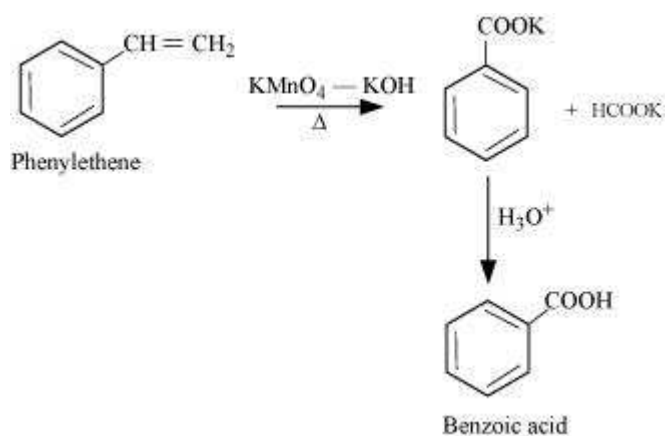
Answer



(iii)



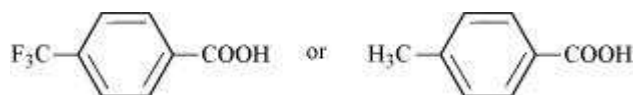
(iv)

**Question 12.8:**

Which acid of each pair shown here would you expect to be stronger?

(i) $\text{CH}_3\text{CO}_2\text{H}$ or $\text{CH}_2\text{FCO}_2\text{H}$ (ii) $\text{CH}_2\text{FCO}_2\text{H}$ or $\text{CH}_2\text{ClCO}_2\text{H}$ (iii) $\text{CH}_2\text{FCH}_2\text{CH}_2\text{CO}_2\text{H}$ or $\text{CH}_3\text{CHFCH}_2\text{CO}_2\text{H}$

(iv)



Answer

(i)



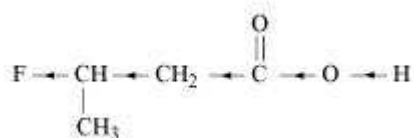
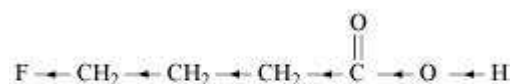
The +I effect of $-\text{CH}_3$ group increases the electron density on the O–H bond. Therefore, release of proton becomes difficult. On the other hand, the $-I$ effect of F decreases the electron density on the O–H bond. Therefore, proton can be released easily. Hence, $\text{CH}_2\text{FCO}_2\text{H}$ is a stronger acid than $\text{CH}_3\text{CO}_2\text{H}$.

(ii)



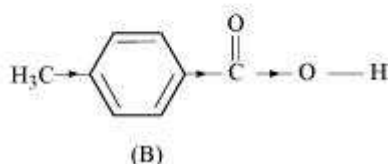
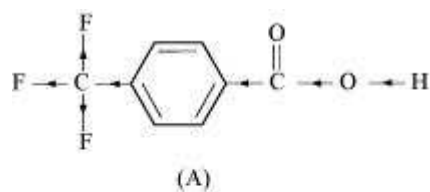
F has stronger $-I$ effect than Cl. Therefore, $\text{CH}_2\text{FCO}_2\text{H}$ can release proton more easily than $\text{CH}_2\text{ClCO}_2\text{H}$. Hence, $\text{CH}_2\text{FCO}_2\text{H}$ is stronger acid than $\text{CH}_2\text{ClCO}_2\text{H}$.

(iii)



Inductive effect decreases with increase in distance. Hence, the +I effect of F in $\text{CH}_3\text{CHFCH}_2\text{CO}_2\text{H}$ is more than it is in $\text{CH}_2\text{FCH}_2\text{CH}_2\text{CO}_2\text{H}$. Hence, $\text{CH}_3\text{CHFCH}_2\text{CO}_2\text{H}$ is stronger acid than $\text{CH}_2\text{FCH}_2\text{CH}_2\text{CO}_2\text{H}$.

(iv)

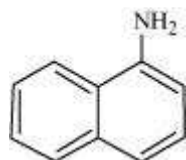


Due to the $-I$ effect of F, it is easier to release proton in the case of compound (A). However, in the case of compound (B), release of proton is difficult due to the +I effect of $-\text{CH}_3$ group. Hence, (A) is a stronger acid than (B).

Question 13.1:

Classify the following amines as primary, secondary or tertiary:

(i)



(ii)



(iii) $(\text{C}_2\text{H}_5)_2\text{CHNH}_2$

(iv) $(\text{C}_2\text{H}_5)_2\text{NH}$

Answer

Primary: **(i)** and **(iii)**

Secondary: **(iv)**

Tertiary: **(ii)**

Question 13.2:

(i) Write structures of different isomeric amines corresponding to the molecular formula, $\text{C}_4\text{H}_{11}\text{N}$

(ii) Write IUPAC names of all the isomers.

(iii) What type of isomerism is exhibited by different pairs of amines?

Answer

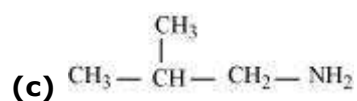
(i), (ii) The structures and their IUPAC names of different isomeric amines corresponding to the molecular formula, $\text{C}_4\text{H}_{11}\text{N}$ are given below:

(a) $\text{CH}_3\text{-CH}_2\text{-CH}_2\text{-CH}_2\text{-NH}_2$

Butanamine (1°)

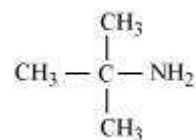
(b) $\text{CH}_3\text{-CH}_2\text{-}\overset{\text{NH}_2}{\underset{|}{\text{CH}}}\text{-CH}_3$

Butan-2-amine (1°)

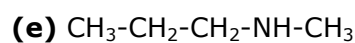


2-Methylpropanamine (1°)

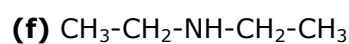
(d)



2-Methylpropan-2-amine (1°)

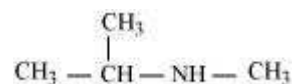


N-Methylpropanamine (2°)



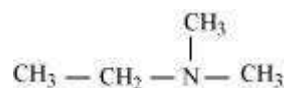
N-Ethylethanamine (2°)

(g)



N-Methylpropan-2-amine (2°)

(h)



N,N-Dimethylethanamine (3°)

(iii) The pairs (a) and (b) and (e) and (g) exhibit position isomerism.

The pairs (a) and (c); (a) and (d); (b) and (c); (b) and (d) exhibit chain isomerism.

The pairs (e) and (f) and (f) and (g) exhibit metamerism.

All primary amines exhibit functional isomerism with secondary and tertiary amines and vice-versa.

Question 13.3:

How will you convert?

(i) Benzene into aniline

(ii) Benzene into N, N-dimethylaniline

(iii) $\text{Cl} - (\text{CH}_2)_4 - \text{Cl}$ into hexan-1, 6-diamine?

Answer
